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Electroosmotic flow and electroosmosis in soil

Tamanna Punia and Attar Singh

Abstract

There is motion of a fluid with respect to a rigid wall when an electric potential gradient is applied across it. This process is called electroosmosis. The aspect by which rate of electro-osmotic flow is managed is the coefficient of electro-osmotic permeability of the soil which is zeta potential dependent. Many researches have measured the variation of the coefficient of electro-osmotic permeability of soils. The value varies from 4.91×10^{-6} to $1.57 \times 10^{-5} \text{ cm}^2 \text{ S}^{-1} \text{ V}^{-1}$. The Helmholtz-Smoluchowski model is largely used to explain electroosmosis process. In this paper first of all the cause of flow is discussed and then electroosmosis in soil is studied.

Keywords: Electroosmosis, Potential gradient, electroosmotic flow, Zeta Potential

Introduction

Electroosmosis is the process of transportation of water through an endless soil particle system, where the motion is firstly produced in the moisture film in the diffuse double layer (where the cations command). Reuss (1809) discovered that water flow could be induced through a capillary by an external electrical gradient. With the application of the straight electrical gradient to a clay-water system, the mobile diffused layer moves and the solution with it is carried while the surface or particle is fixed. The main mechanism in the electroosmosis is the movement of charged particles; anions (-ve charge) move towards +vely charged electrodes (anode) and cations (+ve charge) move towards -vely charged electrode (cathode) (fig. 1).

Mausmann (1990) has proposed higher soil particle surface results in greater moisture film transfers. Viscosity and electrical potential applied can also affect this process [22]. There are many other theories which explain theory of electroosmosis including Helmholtz-Smoluchowski theory, Buckingham TC theory, Ion hydration theory, Spiegler friction model, Schmid theory [6].

The flow which is the result of the fluid surrounding the soil particles is persuaded by ionic fluxes is Electro-Osmotic flow [12]. Also, the direction of flow of the water molecules in bulk phase and flow from fluid surrounding the soil particles is same. Flow in the fluid surrounding the soil particles is first region and flow in bulk phase is second region and when interaction takes place between the two, it enables the motion of water in bulk phase which implies that a drag action is the major reason of the electro-osmotic flow. Therefore, movement of these two water layers is main cause of the overall noticed electro-osmotic flow. Electroosmosis occurs from cathode to anode since positive charged surface has opposing action.

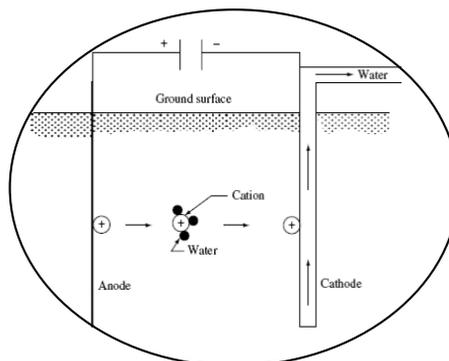


Fig 1: Principles of Electroosmosis (Modified from [21])

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Cause of Electroosmotic flow

Electroosmotic flow takes place when Coulomb force is induced by an electric field on net moving charge in a solution. There is generally a formation of interface (constant), a layer of mobile ions (electrical double layer) formed in the area near the interface due to the chemical equilibrium between any rigid surface and an electrolyte

solution. Net charge in the electrical double layer is induced to move from Coulomb force generated, when an electric field is applied to the fluid. Hence, we get electroosmotic flow. For the determination of electro-osmotic flow, charge distribution in fluid adjacent to flow, soil surface and ζ play significant role according to Hemholtz-Smoluchowski theory.

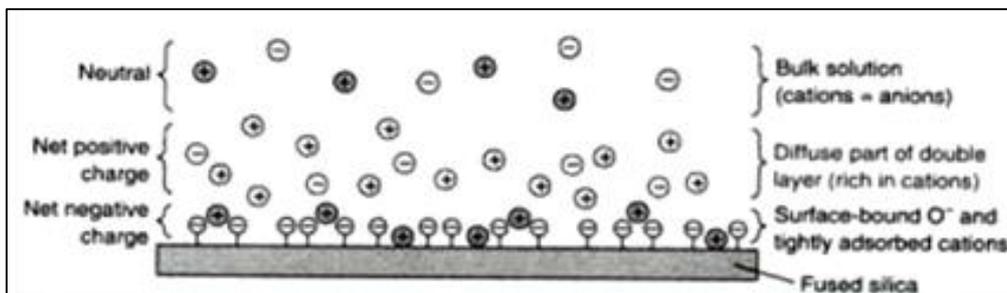


Fig 2: Formation of double layer for electroosmotic flow

Helmholtz-Smoluchowski Theory

To describe electroosmotic process, this theory is widely used. The assumption of the theory is:

(i) The charge distribution in the fluid adjacent to soil surface and the zeta potential play major roles in evaluating electro-osmotic flow. The electric potential developed at liquid-solid interface as a result of the motion of colloidal particles is ζ i.e. the potential at junction between the mobile parts and fixed part of the electrical double layer. Surface potential of particle is more than ζ and lists the value at the slip plane which is placed at a small unknown distance from the colloidal surface [14, 15].

(ii) The size of the pore radii is comparatively big in contrast to the thickness of the diffuse double layer and mobile ions which are collected near the soil-water interface. The value of this potential is also effected by thickness of the double layer, ion exchange capacity and size of ion radius [9].

[10]. Also, sign and magnitude of ζ varies with interaction of both solid and liquid phases.

The coefficient of electro-osmotic permeability of the soil, K_e , which is a measure of the fluid flux per unit area of soil per unit electric gradient regulates the rate of electroosmotic flow. The coefficient of electro-osmotic permeability is given by:-

$$q = \frac{\epsilon \zeta}{V_t} n \frac{E}{L} A$$

where, L = length

ϵ = soil electrical permittivity

q = flow rate.

ζ = zeta potential

A = Gross cross sectional area normal to flow of water

V_t = Viscosity of the pore fluid

n = soil porosity

The value of K_e based on the Helmholtz Smoluchowski theory depends mainly on ζ and n. One of the significant electrokinetic properties of soil colloids is ζ .

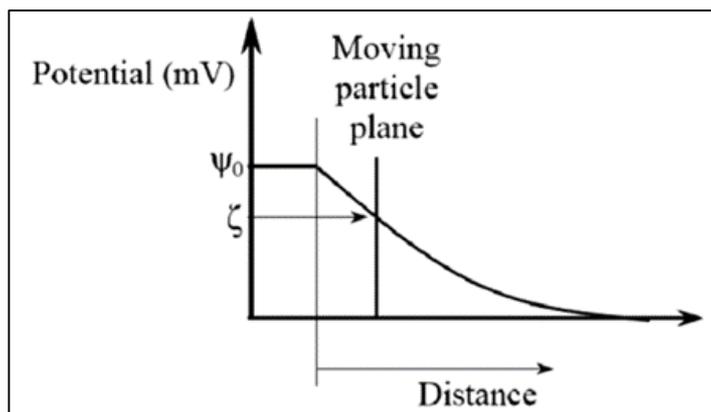


Fig 3: Distribution of potential indicating the slipping plane (Zeta) potential (Modified from [13])

After many researches on ζ potential of clay minerals in various solutions [19, 28], it has been found that valence of ions, concentration of electrolyte, pH and type of electrolyte are significant factors which can affect ζ values.

During electrokinetic process, as long as there is zero change in pH of the pore liquid and the concentration of ions, it is assumed that K_e remains constant. The measurement of ζ is simple and direct and can be measured by Zeta meter. Firstly, the sample is put in a chamber which is known as electrophoresis cell. After that electric field is applied,

consequently particles move with velocity that is directly proportional to Zeta potential. To know whether their change is positive or negative, the direction of the movement of particles is considered.

It is generally preferable to calculate ζ in millivolts instead of in electrostatic units. The formula which shows relation is:

$$\zeta = 11300 V_t \frac{EM}{D_i}$$

which is obtained from most simple expression for ζ ,

$$\zeta = 4\pi V_t \frac{EM}{D_t}$$

where, EM= Electrophoretic mobility at actual temperature

D_t = dielectric constant

ζ = Zeta potential

V_t = Velocity of suspending liquid

As ζ increases the thickness of the double layer increase. The value of ζ is negative for clay. Due to negative ζ , electroosmosis occurs from anode to cathode whereas positive surface charge causes electroosmosis to occur from cathode to anode. There is zero variation of ζ in organic soils while in clay, variation of ζ has been observed. Both ζ and acidity are inversely proportional.

As long as soils with big pores are soaked with water assumption (b) is valid. Helmholtz Smoluchowski equation is less applicable for unsaturated soils or small capillaries.

Casagrande [17] has proposed the most widely used electro-osmotic flow equation for soil system (Eq. 1)

$$q = i_e AK_e \quad (\text{Eq. 1})$$

Where,

K_e = Coefficient of electro-osmotic permeability

q = flow rate

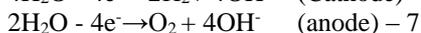
i_e = Applied electrical gradient

A = Gross cross sectional area normal to water flow

Table 1: is showing source general values of K_e for various soils (adjusted from [4, 13, 21])

Soil type	K_e , $\text{Cm}^2/\text{V.S}$
Peat	4.91×10^{-6} to 1.57×10^{-5}
Clay slit	5.0×10^{-5}
Kaolin	5.7×10^{-5}
London clay	5.8×10^{-5}
Boston Blue clay	5.1×10^{-5}
Na- montmorillonite	2.0×10^{-5} to 12×10^{-5}

Electrolysis reactions at electrodes are caused by the application of straight (direct) current through electrodes [9, 10, 24]. Base and Acid are generated by the reduction at the cathode and oxidation of water at the anode respectively. The reactions are:



The acid produced at the anode proceeds through the soil towards cathode due to electroosmosis and ionic migration. Similarly, because of ionic migration and diffusion the base formed at the cathode firstly moves towards the anode. However, the opposite flow due to the electroosmosis makes migration of the base front and back diffusion slower. The movement of the acid front is faster than movement of the base front due to following reasons:

a) Nature of electroosmotic flow is counteracting.

b) H^+ has more ionic mobility than OH^- [26-32].

The electrokinetic process can be increased or decreased, since in geotechnical reactions, soil pores remarkably affect electrokinetic phenomena. The reactions such as dissolution, complexation and precipitation are called geochemical reaction and are highly dependent on pH conditions [5, 18, 32].

Conclusion

Electro-osmosis of soil is the process in which when soil is placed between two electrodes in a fluid. An electromotive force develops and hence the fluid will move from one side to

the other under the application of electric field. Reuss was first to discover that a water flow could be induced through a capillary by an external electric field. A net electric driving force moves the water layer from anode to cathode due to the presence of large amount of positive charges on the surface of soil.

The Zeta potential based on Helmholtz-Smoluchowski plays an important role in electro-osmotic phenomenon.

References

- Asadi A, Moayedi H, Huat BBK, Zamani FBA, Parsaie Sojoudi S. Int. J Electrochem. Sci. 2011; 6:1146-1158.
- Asadi A, Huat BBK, Hanafi MM, Mohamed TA, Shariatmadari N. Geosciences Journal. 2009; 13(2):175-181.
- Asadi A, Huat BBK, Hanafi MM, Mohamed TA, Shariatmadari N. Geosciences Journal. 2010; 14(1):65-75.
- Alshawabkeh AN, Sheahan TC, Wu X. Mechanics of Materials. 2004; 36:453-465.
- British Standard Institution, Methods of test for soils for civil engineering purposes. BSI 1377, Part 1-9, HMSO, London, UK, 1990, 427.
- Gray DH. *Geotechnique*, 1970; 20:81-93.
- Eykholt GR, Daniel DE. Journal of Geotechnical Engineering. ASCE, 1994; 120(5):797-815.
- Helmholtz H. *Annalen der Physik*, 1879; 7:337-382.
- Fang HY, Daniels JL. Introductory Geotechnical Engineering. An environmental perspective. Taylor and Francis, Landon and New York, 2006, 545.
- Fang HY. Introduction to Environmental Geotechnology, CRC Press, Boca Raton, FL, 1997, 652.
- Hamed J, Acar YB, Gale RJ. Journal of Geotechnical Engineering, ASCE, 1991; 117(2):241-271.
- Lyklema J, Minor M. Physicochemical Engineering Aspects, 1998; 140:33-41.
- Mitchell JK, Soga K. Fundamentals of Soil Behavior. John Wiley and Sons, New Jersey, 2005, 577.
- Hunter JR. Introduction to Modern Colloid Science, Oxford Science Publications.
- Hunter JR. Zeta potential in colloid science. Academic Press, London, 1981, 386.
- Hamed JT, Bhadra A. Journal of Hazardous Materials. 1997; 55:279-294.
- Casagrande L. Journal Boston Society of Civil Engineers. 1983; 69(2):255-302.
- Casagrande L. *Geotechnique*, 1949; 1(3):159-177.
- West LJ, Stewart DI, Geotechnical Special Publication No. 46, ASCE, New York, N.Y. 1995; 2:1535-1549.
- West LJ, Stewart DI, Binley AM, Shaw B. Engineering Geology, 1999; 53:205-215.
- Das MB. Advanced Soil Mechanics. Taylor and Francis, New York, 2008, 594.
- Hausmann MR. Engineering Principles of Ground Modifications, New York: McGraw Hill Publishing Company, 1990.
- Mohamedelhassan E, Shang JQ. Canadian Geotechnical Journal. 2003; 40(6):1185-1199.
- Cherepy N, Wildenschild D. Environmental Science and Technology. 2003; 37:3024-3030.
- Lorenz PB. Clays and Clay Minerals, 1969; 17:223-231.
- Probstein RF, Hicks RE. Science, 1993; 260:498-504.
- Smoluchowski VM, Handbuch der Elektrizitat under Magnetismus, 1921; II(2):366-428.
- Yukselen Y, Erzin Y. Environmental Geology, 2008; 54:1059-1066.

29. Acar YB, Alshawabkeh AN. Environmental Science and Technology. 1993; 27(13):2638-2647.
30. Acar YB, Rabbi MF, Ozsu EE. Journal of Geotechnical and Geoenvironmental Engineering. 1997, 239-249.
31. Acar YB, Gale RJ, Putnam GA, Hamed J, Wong RL. Journal of Environmental Science and Health. 1990; A(6):687-714.
32. Acar YB, Gale RJ, Hamed J, Putnam G. Transportation Research Record, Geotechnical Engineering. 1990; 1288:23-34.